

## ions and there bonds

1. Noun
2. Verb
3. Noun
4. Adjective
5. Plural Noun
6. Plural Noun
7. Noun
8. Noun
9. Adjective
10. Noun
11. Noun
12. Noun
13. Adjective

# ions and there bonds

recall that atoms are \_\_\_\_\_<sup>Noun</sup>; they do not have an electric charge. Also \_\_\_\_\_<sup>Verb</sup> that for an \_\_\_\_\_<sup>Noun</sup> to be most \_\_\_\_\_<sup>Adjective</sup>, the outermost energy level should be either empty or completely filled. Some \_\_\_\_\_<sup>Plural noun</sup> tend to give up (donate) or obtain (accept) \_\_\_\_\_<sup>Plural noun</sup> to empty or fill the outer energy level to be stable. An atom that has lost or gained one or more electrons becomes an ion and carries an electric charge. For example, sodium has one electron in its outermost energy level. \_\_\_\_\_<sup>Noun</sup> can become more stable if it gives up this one electron, leaving its outer energy level empty. When it gives away this one negative charge, the neutral \_\_\_\_\_<sup>Noun</sup> atom becomes a \_\_\_\_\_<sup>Adjective</sup> charged \_\_\_\_\_<sup>Noun</sup> ion (Na<sup>+</sup>). Similarly, chlorine has seven electrons in its outer energy level and needs just one electron to fill it. when chlorine accepts an electron from a donor \_\_\_\_\_<sup>Noun</sup> such as \_\_\_\_\_<sup>Noun</sup> chlorine becomes a \_\_\_\_\_<sup>Adjective</sup> charged ion (Cl<sup>-</sup>).