Moles

1. First Name _____
2. Last Name _____
3. Noun _____
4. Noun _____
5. Number _____
6. Noun _____
7. Number _____
8. Noun _____
9. Noun _____
10. Same Noun As Two Ago _____
11. Same Noun As Last _____
12. Number _____
13. Same Number As Last _____
14. Number _____
15. Number _____
16. Noun _____
17. Noun _____

Moles

_____Last name_____was an Italian _____Noun_____who provided us with his number 6.02x1023. This number shows us how many molecules in a mole. Every element has a molar mass, and in turn every compound has a molar mass. For example _____Noun_____molar mass is _____Number_____g/mol but _____Noun_____molar mass is _____Number_____g/mol. The molar mass is the same as the atomic number of the element. You can use the molar mass to find _____Noun_____in _____Noun_____and vice-versa. The way to do _____Same noun as two ago_____to _____Same noun as last_____is by multiplying the molar mass by the amount of moles and to find moles in mass you divide the atoms over molar mass. Molar volume is equal to 22.4L of one mole of an ideal gas at STP. STP is Standard Temperature and (which is 0? C or 32 ? F) and Pressure (which is 1 atm, sea level)> Like molar mass, molar volume can be used to find volume in moles and vice-versa. An example would be having _____Number_____of O2 at STP and you wanted to know how many moles that is. You would divide _____Same number as last_____by _____Number_____to get _____Number_____moles. To get moles into volume you would multiply the amount of moles by 22.4L. Lastly, there is percentage composition. This is when you have a _____Noun_____ and you take the weight of one _____Noun_____and divide it by the other, then you put the decimal into percentage.